

MAP OF MATTER



MATTER

PURE SUBSTANCES

MIXTURES

ELEMENTS

COMPOUNDS

HOMOGENEOUS MIXTURES

HOTEROGENEOUS MIXTURES

:EXAMPLES:

Gold
 Silver
 Hydrogen
 Oxygen
 Iron
 Carbon

:EXAMPLES:

Carbon dioxide
 Water
 Magnesium oxide
 Salt (Sodium chloride)
 Lime (Calcium oxide)

SOLUTIONS

SOLID MIXTURES

COLLOIDS

SUSPENSIONS

:EXAMPLES:

Salt + Water
 Sugar + Water
 Coffee
 Brass
 Air
 Soapy water
 Ink

:EXAMPLES:

Salt + Sand
 Husk + Grain
 Rice + Stones
 Dil Pasand
 Pizza
 Vegetables in bag

:EXAMPLES:

Milk
 Blood
 Fog
 Clouds
 Smoke
 Oil + Water
 Muddy water
 Nail polish

:EXAMPLES:

Dust in Air
 Sandy water
 Starch solution
 Payasam
 Flour + Water

Matter : Anything that has mass and occupies space.

Pure Substances : Either elements or compounds. These can not be divided into smaller parts by any physical methods.

Elements : A substance with homogeneous atoms. These can not be broken down by any physical or chemical methods.

Compounds : A substance with heterogeneous atoms. Formed due to chemical reaction between elements. These can not be broken down into smaller parts by any physical method but can be broken down by chemical methods.

Mixtures : Combination of at least two pure substances. These can be separated into components by physical or chemical methods. (combination of elements or compounds or elements and compounds.)

Homogeneous Mixtures : A mixture having uniform composition throughout.

Heterogeneous mixtures : A mixture not having uniform composition throughout.

Solution : A homogeneous mixture. It has same properties throughout.

Colloids : A heterogeneous mixture. The particles are too small to be seen with the eye but big enough to scatter light. The particles do not settle down. These can not be separated by filtration.

* Emulsion, Gel, Foam, Aerosol, Solid sol, Sol are some types of colloids.

Suspensions : A heterogeneous mixture. The particles are visible to the eye also big enough to scatter light. The particles settle down. These can be separated by filtration.