

SLIP TEST- 9

CHAPTER- 9 : CLASSIFICATION OF ELEMENTS – THE PERIODIC TABLE

Name:..... Section:..... Roll No:..... Max.Marks:20

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

1) How does the following properties change in a period and in a group in periodic table of elements.

(i) Atomic size (ii) Ionization energy

2) Explain the construction of Modern periodic table.

II. Answer the following questions briefly. Each carries two marks. 2 x 2 = 4 M

3) Which group of elements are called halogens? Name them.

4) On the basis of atomic numbers predict to which block the elements with atomic number 37 and 58 belongs to?

III. Answer the following in one or two sentences. Each carries one marks. 2 x 1 = 2 M

5) Which are called representative elements?

6) State Doberienner's law of triads.

IV. Choose the correct choice and write down in the given brackets. 6 x 1 = 6 M

7) Actinides []

A. ${}_{90}\text{Th}$ to ${}_{103}\text{Lr}$ B. ${}_{89}\text{Ac}$ to ${}_{102}\text{No}$ C. ${}_{89}\text{Ac}$ to ${}_{103}\text{Lr}$

D. Any of the above / All of the above

8) Which of the following is the most active metal []

A. Lithium

B. Potassium

C. Sodium

D. Rubidium

9) An element 'X' having configuration 2,8,2 belongs to []

A. 2nd periodB. 2nd group

C. both A and B

D. neither A nor B

10) Identify the Metalloid []

A. Germanium

B. Aluminium

C. Phosphorous

D. Iodine

11) Which of the following is not a Doberienner's triad []

A. Ca,Sr,Ba

B. S,Se,Te

C. Li,Na,K

D. Mn.Co,Fe

12) Number of valence electrons present in elements of group 16 []

A. 2

B. 6

C. 4

D. 8