

SLIP TEST- 10
CHAPTER- 10 : CHEMICAL BONDING

Name:..... Section:..... Roll No:..... Max.Marks:20

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Draw the shape of Ammonia molecule. Explain it. Why the bond angle is $107^{\circ} 48'$?
2) Explain the formation of double bond in Oxygen molecule?

II. Answer the following questions briefly. Each carries two marks. 2 x 2 = 4 M

- 3) Name the molecules having p-p overlap.
4) Draw the shape of water molecule.

III. Answer the following in one or two sentences. Each carries one marks. 2 x 1 = 2 M

- 5) What is the valency of Fluorine?
6) If valency of 'X' is 2 and the valency of 'Y' is 3. Then write the formula of compound that formed due to bond between 'X' and 'Y'.

IV. Choose the correct choice and write down in the given brackets. 6 x 1 = 6 M

- 7) Which of the following element is electro negative? []
A. Sodium B. Oxygen C. Magnesium D. Calcium
- 8) The bond in NaCl molecule is []
A. Covalent B. Ionic C. Polar Covalent D. Coordinate Covalent
- 9) Shape of *Beryllium Chloride* molecule []
A. Triangle B. Pyramidal C. Linear D. Tetrahedral
- 10) Bond length of H-H in H_2 molecule []
A. 0.74 \AA B. 1.44 \AA C. 1.95 \AA D. 1.27 \AA
- 11) Which of the following formed double positive anion []
A. Sodium B. Magnesium C. Oxygen D. Chlorine
- 12) sigma bonds present in Ethylene C_2H_4 []
A. 2 B. 4 C. 1 D. 3