1. **Class 10**

1. Class 10th students should complete the assignment by the due date. The class teacher will discuss the topic in detail during the next class. Please make sure to complete the assignments and submit them on time.

2. **Klassenliste**

   1. A list of students is available in the classroom. Please check your name on the list. If you are not present, please inform the class teacher.

3. **Mathematics Class**

   1. The mathematics class will start at 9 AM sharp. Please be on time.

4. **Science Class**

   1. The science class will start at 10 AM sharp. Please be on time.

5. **Physics Class**

   1. The physics class will start at 11 AM sharp. Please be on time.

6. **Chemistry Class**

   1. The chemistry class will start at 12 PM sharp. Please be on time.

7. **Biotechnology Class**

   1. The biotechnology class will start at 1 PM sharp. Please be on time.

2. **Chemistry Class: Compound Formation**

   1. **1) 3NaBr + H₃PO₄ → Na₃PO₄ + 3 HBr**

   2. **2) 3Mg + Fe₂O₃ → 2 Fe + 3 MgO**

   3. **3) C₂H₄ + 3O₂ → 2 CO₂ + 2 H₂O**

   4. **4) 2 PbSO₄ → 2PbSO₃ + O₂**

   5. **5) 2NH₃ + 3I₂ → N₂H₆ + 3H₂**

   6. **6) H₂O + SO₃ → H₂SO₄**

3. **Chemistry Class: Compound Formation**

   1. **1) NaNO₃ + PbO → Pb(NO₃)₂ + Na₂O**

   2. **2) AgI + Fe₂(CO₃)₂ → Fe₃ + Ag₂CO₃**

   3. **3) C₂H₄O₂ + O₂ → CO₂ + H₂O**

   4. **4) ZnSO₄ + Li₂CO₃ → ZnCO₃ + Li₂SO₄**

   5. **5) V₂O₅ + CaS → CaO + V₂S₅**

   6. **6) Mn(NO₂)₂ + BeCl₂ → Be(NO₂)₂ + MnCl₂**

   7. **7) AgBr + GaPO₄ → Ag₃PO₄ + GaBr₃**

   8. **8) H₂SO₄ + B(OH)₃ → B₂(SO₄)₃ + H₂O**

   9. **9) S₈ + O₂ → SO₂**

   10. **10) Fe + AgNO₃ → Fe(NO₃)₂ + Ag**

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3. Class-10 Solved Questions

1. How many players are needed to play the game?
   - Option (a) 2 players
   - Option (b) 3 players
   - Option (c) 4 players
   - Option (d) 5 players
   - Option (e) 6 players
   - Option (f) 7 players

2. What is the name of the game?
   - Option (a) Cricket
   - Option (b) Football
   - Option (c) Badminton
   - Option (d) Tennis
   - Option (e) Hockey
   - Option (f) Basketball

3. How many goals are scored in a single game?
   - Option (a) 1 goal
   - Option (b) 2 goals
   - Option (c) 3 goals
   - Option (d) 4 goals
   - Option (e) 5 goals
   - Option (f) 6 goals

4. What is the name of the coach who is coaching the team?
   - Option (a) Mr. Singh
   - Option (b) Mr. Johnson
   - Option (c) Ms. Smith
   - Option (d) Ms. Jones
   - Option (e) Mr. Brown
   - Option (f) Ms. Green

5. What is the name of the player who scored the most goals?
   - Option (a) John
   - Option (b) Jane
   - Option (c) Mike
   - Option (d) Sarah
   - Option (e) Tom
   - Option (f) Emily

6. How many players are in the starting lineup?
   - Option (a) 5 players
   - Option (b) 6 players
   - Option (c) 7 players
   - Option (d) 8 players
   - Option (e) 9 players
   - Option (f) 10 players

7. What is the name of the team?
   - Option (a) RED
   - Option (b) BLUE
   - Option (c) GREEN
   - Option (d) YELLOW
   - Option (e) WHITE
   - Option (f) ORANGE

8. How many games are played in a season?
   - Option (a) 10 games
   - Option (b) 20 games
   - Option (c) 30 games
   - Option (d) 40 games
   - Option (e) 50 games
   - Option (f) 60 games

9. What is the name of the coach's assistant?
   - Option (a) Mr. Smith
   - Option (b) Ms. Johnson
   - Option (c) Mr. Brown
   - Option (d) Ms. Green
   - Option (e) Mr. White
   - Option (f) Ms. Black

4. Class-11 Solved Questions

1. How many questions are in the test?
   - Option (a) 10 questions
   - Option (b) 20 questions
   - Option (c) 30 questions
   - Option (d) 40 questions
   - Option (e) 50 questions
   - Option (f) 60 questions

2. What is the name of the subject?
   - Option (a) Physics
   - Option (b) Chemistry
   - Option (c) Biology
   - Option (d) Mathematics
   - Option (e) English
   - Option (f) History

3. How many marks are allocated for the test?
   - Option (a) 100 marks
   - Option (b) 200 marks
   - Option (c) 300 marks
   - Option (d) 400 marks
   - Option (e) 500 marks
   - Option (f) 600 marks

4. What is the name of the professor who is teaching the subject?
   - Option (a) Dr. Singh
   - Option (b) Dr. Johnson
   - Option (c) Dr. Smith
   - Option (d) Dr. Jones
   - Option (e) Dr. Brown
   - Option (f) Dr. Green

5. How many students are in the class?
   - Option (a) 10 students
   - Option (b) 20 students
   - Option (c) 30 students
   - Option (d) 40 students
   - Option (e) 50 students
   - Option (f) 60 students

6. What is the name of the book used for the test?
   - Option (a) Physics for Class-11
   - Option (b) Chemistry for Class-11
   - Option (c) Biology for Class-11
   - Option (d) Mathematics for Class-11
   - Option (e) English for Class-11
   - Option (f) History for Class-11

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5. Multiple Choice Questions

1. (a) Conduction (b) Convection (c) Radiation (d) Refraction

2. (a) Water (b) Earth (c) Air (d) Fire

3. (a) Conduction (b) Convection (c) Radiation (d) Refraction

4. (a) Conduction (b) Convection (c) Radiation (d) Refraction

5. (a) Conduction (b) Convection (c) Radiation (d) Refraction

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6. *

1. \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है जो हालांकि विरोधी है। जैसा कि \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
2. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
3. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
4. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।

(a) \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है।
(b) \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है।
(c) \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है।
(d) \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है।
(e) \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है।
(f) \( \text{C}_2 \text{H}_2 \) अरू रूपान्तरित किया गया है।

5. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
6. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
7. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
8. \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण अलग से लिखा गया है। \( \text{C}_2 \text{H}_2 \) का नाम लेने वाले अन्य \( \text{C}_2 \text{H}_2 \) नामक अंगरक्षण सम्यक तथ्य के आधार पर ही है।
8. Notes on Elements

1. Chromium, nickel, and copper are metals that are used in electronics and construction.
2. Magnesium and calcium are alkaline earth metals and are used in alloys.
3. Sodium and potassium are alkali metals and are used in batteries.
4. Iron and aluminum are transition metals and are used in construction.
5. S, P, and d electrons - An electron is a particle that is around an atom.
6. Alkali metals are used in batteries.
7. Transition metals are used in construction.
8. Alkaline earth metals are used in alloys.

9. More Notes on Elements - Carbon Family

1. VIII A (carbon family) is a group of elements that react with oxygen to form carbon dioxide.
2. IA (aluminum family) is a group of elements that react with oxygen to form aluminum oxide.
3. Carbon (carbon family) is a group of elements that react with oxygen to form carbon dioxide.
4. Nitrogen (nitrogen family) is a group of elements that react with oxygen to form nitrogen oxide.
5. Oxygen (oxygen family) is a group of elements that react with oxygen to form oxygen gas.
6. Fluorine (fluorine family) is a group of elements that react with oxygen to form fluorine gas.
7. Chlorine (chlorine family) is a group of elements that react with oxygen to form chlorine gas.
8. Bromine (bromine family) is a group of elements that react with oxygen to form bromine gas.
9. Iodine (iodine family) is a group of elements that react with oxygen to form iodine gas.
10. Astatine (astatine family) is a group of elements that react with oxygen to form astatine gas.

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10. Three Marks

1. Make a simple pendulum of length 0.6m. Make a wire and take a ruler to measure it. 
   As the length of the pendulum, the time period is taken.

2. Make a simple pendulum of length 0.4m. Make a wire and take a ruler to measure it. 
   As the length of the pendulum, the time period is taken.

3. Make a simple pendulum of length 0.3m. Make a wire and take a ruler to measure it. 
   As the length of the pendulum, the time period is taken.

4. Make a simple pendulum of length 0.2m. Make a wire and take a ruler to measure it. 
   As the length of the pendulum, the time period is taken.

5. Make a simple pendulum of length 0.1m. Make a wire and take a ruler to measure it. 
   As the length of the pendulum, the time period is taken.

(a) Telephone (b) Villa (c) Hotel
(d) Capsule (e) Hotel (f) Flat

11. Two Marks

1. How much energy is stored in a capacitor of 10μF when the potential difference is 2V? 
   How much energy is stored in a capacitor of 20μF when the potential difference is 4V?

2. A capacitor has a capacitance of 1μF and is charged to a potential difference of 12V. 
   How much energy is stored in the capacitor? 
   If 30 joules of energy is used, how much charge is removed from the capacitor?

3. A capacitor has a capacitance of 3μF and is charged to a potential difference of 10V. 
   How much energy is stored in the capacitor? 
   If 5 joules of energy is used, how much charge is removed from the capacitor?

4. A capacitor has a capacitance of 4μF and is charged to a potential difference of 15V. 
   How much energy is stored in the capacitor? 
   If 25 joules of energy is used, how much charge is removed from the capacitor?

5. A capacitor has a capacitance of 5μF and is charged to a potential difference of 20V. 
   How much energy is stored in the capacitor? 
   If 35 joules of energy is used, how much charge is removed from the capacitor?

6. A capacitor has a capacitance of 6μF and is charged to a potential difference of 25V. 
   How much energy is stored in the capacitor? 
   If 50 joules of energy is used, how much charge is removed from the capacitor?

12. One Mark

1. Define the term Lenz's law. 
   How does it relate to the flow of current in a conductor?

2. Explain the principle of magnetic induction. 
   How does it relate to the flow of current in a conductor?

3. Define the term magnetic field. 
   How does it relate to the flow of current in a conductor?

4. Explain the principle of magnetic induction. 
   How does it relate to the flow of current in a conductor?

5. Explain the term magnetic field. 
   How does it relate to the flow of current in a conductor?

6. Define the term magnetic field. 
   How does it relate to the flow of current in a conductor?
7. a. (iii) Explain the concept of insulators. (2 marks)

13. Name the elements

<table>
<thead>
<tr>
<th>Element</th>
<th>Symbol</th>
<th>Molar Mass</th>
<th>Density</th>
</tr>
</thead>
</table>

3. Describe the properties of metals and non-metals. (2 marks)

4. Outline the key differences between metals and non-metals. (2 marks)

5. What are the characteristics of metals and non-metals? (2 marks)

6. Illustrate the differences between metals and non-metals. (2 marks)

14. Describe the nature of the universe

1. Explain the concept of the universe and its components. (2 marks)

2. Discuss the role of the universe in our understanding of the cosmos. (2 marks)

3. Analyze the implications of the universe on our daily lives. (2 marks)

4. Evaluate the impact of the universe on our technology. (2 marks)

5. Reflect on the significance of the universe in our perception of reality. (2 marks)

6. Demonstrate the relevance of the universe to our understanding of the cosmos. (2 marks)