

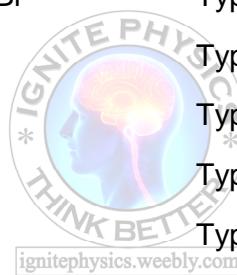
PROJECT WORK-2
CHEMICAL REACTIONS AND EQUATIONS

Write the word equations for each of the following chemical reactions:

- 1) When dissolved beryllium chloride reacts with dissolved silver nitrate in water, aqueous beryllium nitrate and silver chloride powder are made.
- 2) When dissolved sodium hydroxide reacts with sulfuric acid (H_2SO_4), aqueous sodium sulfate, water, and heat are formed.
- 3) When fluorine gas is put into contact with calcium metal at high temperatures, calcium fluoride powder is created in an exothermic reaction.
- 4) When sodium metal reacts with iron (II) chloride, iron metal and sodium chloride are formed.

Observe the following equations and indicate the type of reaction taking place:

- 1) $3NaBr + H_3PO_4 \rightarrow Na_3PO_4 + 3 HBr$ Type of reaction: _____
- 2) $3Mg + Fe_2O_3 \rightarrow 2 Fe + 3 MgO$ Type of reaction: _____
- 3) $C_2H_4 + 3O_2 \rightarrow 2 CO_2 + 2 H_2O$ Type of reaction: _____
- 4) $2 PbSO_4 \rightarrow 2PbSO_3 + O_2$ Type of reaction: _____
- 5) $2NH_3 + 3I_2 \rightarrow N_2I_6 + 3H_2$ Type of reaction: _____
- 6) $H_2O + SO_3 \rightarrow H_2SO_4$ Type of reaction: _____



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Balance the following equations:

- 1) $\underline{\hspace{1cm}} NaNO_3 + \underline{\hspace{1cm}} PbO \rightarrow \underline{\hspace{1cm}} Pb(NO_3)_2 + \underline{\hspace{1cm}} Na_2O$
- 2) $\underline{\hspace{1cm}} AgI + \underline{\hspace{1cm}} Fe_2(CO_3)_3 \rightarrow \underline{\hspace{1cm}} FeI_3 + \underline{\hspace{1cm}} Ag_2CO_3$
- 3) $\underline{\hspace{1cm}} C_2H_4O_2 + \underline{\hspace{1cm}} O_2 \rightarrow \underline{\hspace{1cm}} CO_2 + \underline{\hspace{1cm}} H_2O$
- 4) $\underline{\hspace{1cm}} ZnSO_4 + \underline{\hspace{1cm}} Li_2CO_3 \rightarrow \underline{\hspace{1cm}} ZnCO_3 + \underline{\hspace{1cm}} Li_2SO_4$
- 5) $\underline{\hspace{1cm}} V_2O_5 + \underline{\hspace{1cm}} CaS \rightarrow \underline{\hspace{1cm}} CaO + \underline{\hspace{1cm}} V_2S_5$
- 6) $\underline{\hspace{1cm}} Mn(NO_2)_2 + \underline{\hspace{1cm}} BeCl_2 \rightarrow \underline{\hspace{1cm}} Be(NO_2)_2 + \underline{\hspace{1cm}} MnCl_2$
- 7) $\underline{\hspace{1cm}} AgBr + \underline{\hspace{1cm}} GaPO_4 \rightarrow \underline{\hspace{1cm}} Ag_3PO_4 + \underline{\hspace{1cm}} GaBr_3$
- 8) $\underline{\hspace{1cm}} H_2SO_4 + \underline{\hspace{1cm}} B(OH)_3 \rightarrow \underline{\hspace{1cm}} B_2(SO_4)_3 + \underline{\hspace{1cm}} H_2O$
- 9) $\underline{\hspace{1cm}} S_8 + \underline{\hspace{1cm}} O_2 \rightarrow \underline{\hspace{1cm}} SO_2$
- 10) $\underline{\hspace{1cm}} Fe + \underline{\hspace{1cm}} AgNO_3 \rightarrow \underline{\hspace{1cm}} Fe(NO_3)_2 + \underline{\hspace{1cm}} Ag$