

FORMATIVE ASSESSMENT-3
CHAPTERS - 8,9,10

Name:..... Section:..... Roll No:..... Max.Marks:20

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Define ionization energy. What are the affecting factors of ionization energy? Explain.
2) Write a brief notes about Quantum numbers.

II. Answer the following questions briefly. Each carries two marks. 2 x 2 = 4 M

- 3) Explain the formation of CaO molecule.
4) How the Electro negativity varies in a period and in a group in periodic table of elements.

III. Answer the following in one or two sentences. Each carries one marks. 2 x 1 = 2 M

- 5) How many maximum number of electrons can be accommodated in all d- orbitals in M-shell?

- 6) Name two molecules having double bond.

IV. Choose the correct choice and write down in the given brackets. 6 x 1 = 6 M

- 7) The element having greatest value of Electro Negativity []

A. Fluorine B. Chlorine C. Lithium D. Sodium

- 8) The element having atomic number 48 belongs to block in periodic table. []

A. s-block B. p-block C. d-block D. f-block

- 9) Identify the covalent compound from the following []

A. MgCl₂ B. BeCl₂ C. NaCl D. AlCl₃

- 10) The element having greatest value of Electron affinity []

A. Fluorine B. Chlorine C. Lithium D. Sodium

- 11) Valence electronic configuration of *Chromium* is []

A. 4s²3d⁴ B. 4s¹3d⁵ C. 4s²3d⁹ D. 4s¹3d¹⁰

- 12) Correct method of filling electrons in 1s, 2s, 2p orbitals. []

A.

↑↓	↑↓	↑↑	↑↑	↑↑
----	----	----	----	----

 B.

↑↓	↑↓	↑↓	↑↑	↑↑
----	----	----	----	----

C.

↑↓	↑↓	↑↓	↓↑	↑↑
----	----	----	----	----

 D.

↑↓	↑↓	↑↓	↑↑	↑↑
----	----	----	----	----