

FORMATIVE ASSESSMENT-3
CHAPTERS – 8,9,10

Name:..... Section:..... Roll No:..... Max.Marks:25

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Define ionization energy. What are the affecting factors of ionization energy? Explain.
- 2) Write a brief notes about Quantum numbers.

II. Answer the following questions briefly. Each carries two marks. 3 x 2 = 6 M

- 3) Draw Moeller's chart of showing ascending order of energies of various atomic orbitals.
- 4) Explain the formation of CaO molecule.
- 5) How the Electro negativity varies in a period and in a group in periodic table of elements.

III. Answer the following in one or two sentences. Each carries one marks. 3 x 1 = 3 M

- 6) How many maximum number of electrons can be accommodated in all d- orbitals in M-shell?
- 7) State Modern periodic law.
- 8) Name two molecules having double bond.

IV. Choose the correct choice and write down in the given brackets. 4 x 1 = 4 M

- 9) Identify the covalent compound from the following []
 A. MgCl₂ B. BeCl₂ C. NaCl D. AlCl₃
- 10) The element having greatest value of Electron affinity []
 A. Fluorine B. Chlorine C. Lithium D. Sodium
- 11) Valence electronic configuration of Chromium is * []
 A. 4s²3d⁴ B. 4s¹3d⁵ C. 4s²3d⁹ D. 4s¹3d¹⁰
- 12) Correct method of filling electrons in 1s, 2s, 2p orbitals. []

A.

↑↓	↑↓	↑↑	↑↑	↑↑
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 B.

↑↓	↑↓	↑↓	↑↑	
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C.

↑↓	↑↓	↑↓	↓↑	↑↑
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 D.

↑↓	↑↓	↑↓	↑↑	↑
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IV. Fill in the blanks with suitable answers. 4 x 1 = 4 M

- 13) has the greatest value of Electro negativity.
- 14) The element having atomic number 48 belongs to block in periodic table.
- 15) proposed by ionic bond.
- 16) In nl^x method; indicates principal energy level.

NAGA MURTHY- 9441786635
 Contact at : nagamurthysir@gmail.com
 Visit at : ignitephysics.weebly.com