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**FORMATIVE ASSESSMENT-3**

**Physical Science**

**10<sup>th</sup> Class**

**CCE Model**

**Key to FA-3**

**(2016-17)**

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- (a) Group number in Arabic numerical = **16**.

(b) These elements belong to **p-block**.

(c) The special name of oxygen family is **Chalcogen family**.

(d) The general outer shell configuration of these elements is  **$ns^2np^4$** .
  - (a) The atomic number = **10**.

(b) The element name: **Neon (Ne)**.

(c) The electronic configuration of Ne is  **$1s^22s^22p^6$**  or **2,8**.

(d) Neon (Ne) belongs to **VIII-A group or 18<sup>th</sup> group** and **2<sup>nd</sup> period**.
  - Yes, I support this statement.

Orbit is the path that an electron revolves round the nucleolus.

Orbital is the place at where the probability to find the electron.

So, these two are different.
  - The ionization energy in groups decreases because the atomic size of the atoms increases from top to bottom in the modern periodic table and the attraction between the nucleolus and electron in the outermost shell is decreases. So, it is easy to remove the electron from the top to bottom atoms in the modern periodic table.
  - He used **Hund's Rule**.
  - The elements in the periodic table arranged in the increasing order of their atomic numbers or in the modern Periodic law, i.e., the physical and chemical properties of elements is the periodic functions of the electronic Configurations of their atoms.
7. **b**            8. **c**            9. **b**            10. **b**            11. **d**            12. **b**
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