FORMATIVE ASSESSMENT-1
CHAPTERS - 1,2

Name:………………………………….. Section:………… Roll No:……….....   Max.Marks:20

....

I. Answer the following questions. Each carries four marks.   **2 x 4 = 8 M**

1) The rate of raise in temperature depends upon the nature of substances. How can you prove experimentally?

2) Balance the following chemical equations.

    (i) $AgCl_2$ $\rightarrow$ Ag + Cl_2

    (ii) Zn + HCl $\rightarrow$ $ZnCl_2$ + H_2

    (iii) PbO + C $\rightarrow$ Pb + CO_2

    (iv) CH_4 + O_2 $\rightarrow$ CO_2 + H_2O

II. Answer the following questions briefly. Each carries two marks.   **2 x 2 = 4 M**

3) Write the differences evaporation and boiling.

4) Write any two chemical equations of reactions in which gases are liberated.

III. Answer the following in one or two sentences. Each carries one marks.   **2 x 1 = 2 M**

5) If glucose powder is added to a glass of water. What happened? Identify whether it is an exothermic reaction or an endothermic reaction?

6) Define specific heat.

IV. Choose the correct choice and write down in the given brackets.   **6 x 1 = 6 M**

7) The brown coloured gas released when Lead nitrate is heated is ………   [   ]

    A. Hydrogen   **B.** Nitrogen

    C. Nitrogen dioxide   **D.** Carbon dioxide

8) Boiling point of water is ………………… the boiling point of Spirit.   [   ]

    A. greater than   **B.** less than

    C. equal to   **D.** None of the above

9) 1 calorie = ……………………Joule   [   ]

    A. 4.4   **B.** 4.1   **C.** 4.2   **D.** 4.3

10) Heat flows from   [   ]

    A. Hot body to hot body   **B.** Hot body to cold body

    C. Cold body to hot body   **D.** Cold body to cold body

11) The energy released in which 1 gm of water at $0^{o}C$ freezes to 1 gm of ice at $0^{o}C$ [   ]

    A. 80 cal   **B.** 540 cal   **C.** 640 cal   **D.** 720 cal

12) Precipitate can be shown as …………. In chemical equations.   [   ]

    A. $\leftarrow$ **B.** $\uparrow$ **C.** $\downarrow$ **D.** $\rightarrow$

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