

SLIP TEST-2 (2)

CHAPTER-2 : CHEMICAL REACTIONS AND EQUATIONS

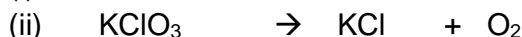
Name:..... Section:..... Roll No:..... Max.Marks:20

I. Answer the following questions. Each carries four marks. 2 x 4 = 8 M

- 1) Draw a neat labeled diagram that shows electrolysis process of water.
- 2) Take two beakers and prepare Sodium sulphate aqueous solution and Barium chloride aqueous solutions. What are the colours of the solutions. Now mix them in another beaker. What happens? What type of chemical reaction it is? What are the products?

II. Answer the following questions briefly. Each carries two marks. 2 x 2 = 4 M

- 3) Balance the following chemical equations.



- 4) Write a brief notes on rancidity.

III. Answer the following in one or two sentences. Each carries one marks. 2 x 1 = 2 M

- 5) What are the components in "Stain less steel"?
- 6) How can we identify Oxygen gas in a laboratory?

IV. Choose the correct choice and write down in the given brackets. 6 x 1 = 6 M

- 7) Not a chemical reaction []

A. Milk turns to curd

B. Burning of Coal

C. Melting of Ice

D. Explosion of crackers

- 8) Colour of Lead iodide []

A. White

B. Yellow

C. Brown

D. Black

- 9) How much Oxygen is needed to burn 12 grams of Carbon []

A. 8 g

B. 16 g

C. 32 g

D. 64 g

- 10) Oxidation []

A. Adding Oxygen

B. Removing Hydrogen

C. Loss of electrons

D. All of the above

- 11) Patina of Copper is in Colour.(Tarnishing of copper) []

A. Red

B. Green

C. Black

D. Brown

- 12) Adding Lime in to water []

A. Exothermic reaction

B. Endothermic reaction

C. Chemical decomposition reaction

D. Precipitate reaction