
FORMATIVE ASSESSMENT-3
PHYSICAL SCIENCE-5th, 6th Chapters
9th Class
CCE Model

Time: 45 min.

2016-17

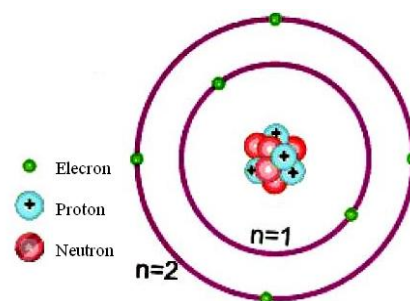
Max. Marks: 20.

Name : _____

Sec : _____ Roll No : _____

I. Answer the following Long answer type questions.

4M X 2 = 8



1. Observe the side given figure, which shows the arrangement of electrons in a neutral atom. Answer the following questions.

- What is the atomic number of this atom?
 - Which atom has this type of arrangement of electrons?
 - Which shells this atom has?
 - What is the valency of this atom?
2. "Matter is neither created nor destroyed during a chemical reaction."
Answer the following questions by using the above statement.
- What is this law?
 - Who established this law?
 - Which postulate of Dalton's theory is the result of this law?
 - Who verified this law experimentally?

II. Answer the following Short answer type questions.

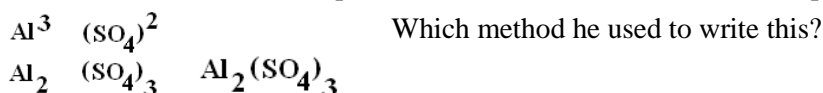
2M X 2 = 4

- "H₂ is different from 2H." Are you support this statement. Explain?
- How ¹²C ¹³C ¹⁴C are called? What is the main difference of these?

III. Answer the following Very Short answer type questions.

1M X 2 = 2

5. Your friend follows these steps to write the formula of aluminium sulphate as given below.



6. Write the relation between Atomic mass number, Atomic number and Neutron number.

**Prepared By: GAMPARAMAPRASAD, PGT (PS) CRT, SRIKAKULAM, Mobile: 7799884688,
Email: gamparamaprasad@gmail.com**

IV. Choose the correct answer of the following.
1M X 6 = 6

7. Which of the following statements are correct in the case of an atom? (c)

- (i) Protons and electrons are called nucleons.
 (ii) Protons, Neutrons and Electrons are called subatomic particles.
 (iii) All the positively charged material in an atom formed a small dense centre, called the nucleus of the atom.
 (iv) Electrons moved around the nucleus of atom in electron shells.
- (a) (i),(ii),(iii),(vi) (b) (i),(ii),(iii) (c) (ii), (iii),(iv) (d) (i), (ii),(iv)

8. Which of the following is not a correct set of atomicity? (d)

- (a) Helium: Monoatomic (b) Hydrogen: Diatomic
 (c) Ozone: Triatomic (d) Sulphur: Tetratomic

9. Match the elements given in Column I with their other names given in Column II.

(c)

Column I	Column II
A. Lead	1. Kalium
B. Tungsten	2. Plumbum
C. Potassium	3. Wolfram

Codes

- | | | | | | | | |
|-----|---|---|---|-----|---|---|---|
| | A | B | C | | A | B | C |
| (a) | 3 | 2 | 1 | (b) | 3 | 1 | 2 |
| (c) | 2 | 3 | 1 | (d) | 2 | 1 | 3 |

10. Ascending order of the molecular mass of H_2O , NaCl and H_2SO_4 molecules is (a)

- (a) $\text{H}_2\text{O} < \text{NaCl} < \text{H}_2\text{SO}_4$ (b) $\text{H}_2\text{SO}_4 < \text{H}_2\text{O} < \text{NaCl}$
 (c) $\text{NaCl} < \text{H}_2\text{O} < \text{H}_2\text{SO}_4$ (d) $\text{H}_2\text{O} < \text{H}_2\text{SO}_4 < \text{NaCl}$

11. Arrange these in an order of discovers or developments of atomic theories/models? (d)

- (i) Nuclear model of an atom
 (ii) Plum pudding model
 (iii) Bohr's model of the atom
 (iv) Dalton's atomic theory.
- (a) (iv), (iii), (ii) and (i) (b) (iv), (i), (ii) and (iii)
 (c) (iv), (ii), (iii) and (i) (d) (iv), (ii), (i) and (iii)

12. The maximum number of electrons in L shell are (b)

- (a) 2 (b) 8 (c) 18 (d) 32
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