

FORMATIVE ASSESSMENT-3  
PHYSICAL SCIENCE-5<sup>th</sup>, 6<sup>th</sup> Chapters  
9<sup>th</sup> Class  
CCE Model

Time: 45 min.

2016-17

Max. Marks: 20.

Name : \_\_\_\_\_

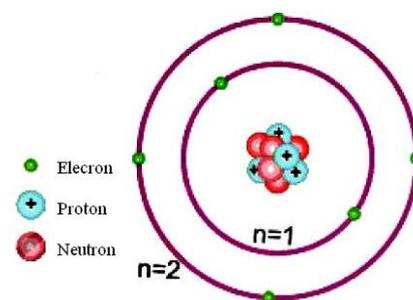
Sec : \_\_\_\_\_ Roll No : \_\_\_\_\_

**I. Answer the following Long answer type questions.**

4M X 2 = 8

1. Observe the side given figure, which shows the arrangement of electrons in a neutral atom. Answer the following questions.

- What is the atomic number of this atom?
- Which atom has this type of arrangement of electrons?
- Which shells this atom has?
- What is the valency of this atom?



2. "Matter is neither created nor destroyed during a chemical reaction."

Answer the following questions by using the above statement.

- What is this law?
- Who established this law?
- Which postulate of Dalton's theory is the result of this law?
- Who verified this law experimentally?

**II. Answer the following Short answer type questions.**

2M X 2 = 4

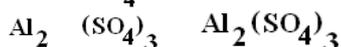
3. "H<sub>2</sub> is different from 2H." Are you support this statement. Explain?

4. How <sup>12</sup>C <sup>13</sup>C <sup>14</sup>C are called? What is the main difference of these?

**III. Answer the following Very Short answer type questions.**

1M X 2 = 2

5. Your friend follows these steps to write the formula of aluminium sulphate as given below.



6. Write the relation between Atomic mass number, Atomic number and Neutron number.

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**IV. Choose the correct answer of the following.**
**1M X 6 = 6**

7. Which of the following statements are correct in the case of an atom? ( c )

- (i) Protons and electrons are called nucleons.  
 (ii) Protons, Neutrons and Electrons are called subatomic particles.  
 (iii) All the positively charged material in an atom formed a small dense centre, called the nucleus of the atom.  
 (iv) Electrons moved around the nucleus of atom in electron shells.
- (a) (i),(ii),(iii),(vi)      (b) (i),(ii),(iii)      (c) (ii), (iii),(iv)      (d) (i), (ii),(iv)

8. Which of the following is not a correct set of atomicity? ( d )

- (a) Helium: Monoatomic      (b) Hydrogen: Diatomic  
 (c) Ozone: Triatomic      (d) Sulphur: Tetratomic

9. Match the elements given in Column I with their other names given in Column II.

( c )

| Column I     | Column II  |
|--------------|------------|
| A. Lead      | 1. Kalium  |
| B. Tungsten  | 2. Plumbum |
| C. Potassium | 3. Wolfram |

Codes

- |     |   |   |   |     |   |   |   |
|-----|---|---|---|-----|---|---|---|
|     | A | B | C |     | A | B | C |
| (a) | 3 | 2 | 1 | (b) | 3 | 1 | 2 |
| (c) | 2 | 3 | 1 | (d) | 2 | 1 | 3 |

10. Ascending order of the molecular mass of  $\text{H}_2\text{O}$ ,  $\text{NaCl}$  and  $\text{H}_2\text{SO}_4$  molecules is ( a )

- (a)  $\text{H}_2\text{O} < \text{NaCl} < \text{H}_2\text{SO}_4$       (b)  $\text{H}_2\text{SO}_4 < \text{H}_2\text{O} < \text{NaCl}$   
 (c)  $\text{NaCl} < \text{H}_2\text{O} < \text{H}_2\text{SO}_4$       (d)  $\text{H}_2\text{O} < \text{H}_2\text{SO}_4 < \text{NaCl}$

11. Arrange these in an order of discovers or developments of atomic theories/models? ( d )

- (i) Nuclear model of an atom  
 (ii) Plum pudding model  
 (iii) Bohr's model of the atom  
 (iv) Dalton's atomic theory.
- (a) (iv), (iii), (ii) and (i)      (b) (iv), (i), (ii) and (iii)  
 (c) (iv), (ii), (iii) and (i)      (d) (iv), (ii), (i) and (iii)

12. The maximum number of electrons in L shell are ( b )

- (a) 2      (b) 8      (c) 18      (d) 32
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